

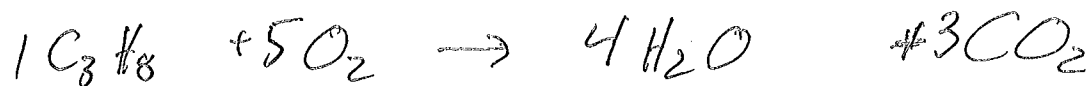
Practice

Write a complete, balanced chemical equation for the combustion of each of these fuels. Include subscripts as needed.

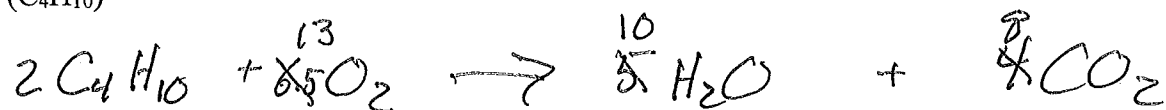
1. Methanol (CH_4O)



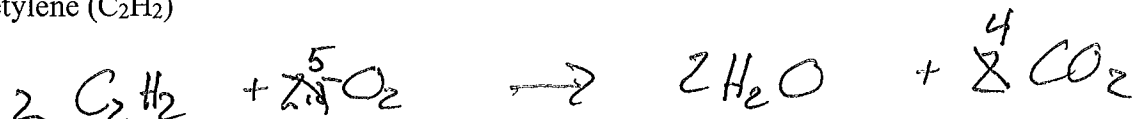
2. Propane (C_3H_8)



3. Butane (C_4H_{10})



4. Acetylene (C_2H_2)



5. Ethanol ($\text{C}_2\text{H}_6\text{O}$)



6. Gasoline – Hexane (C_6H_{14})



Write a complete, balanced chemical equation for the incomplete combustion of each of these fuels. Include subscripts as needed.

7. Gasoline - Octane (C_8H_{18}) $\text{C}_8\text{H}_{18} + 11\text{O}_2 \rightarrow 6\text{CO}_2 + 9\text{H}_2\text{O} + 1\text{CO} + 1\text{C}$

8. Glucose ($\text{C}_6\text{H}_{12}\text{O}_6$)

